TABLE 1--Amounts of calcite, anhydrite, and halite obtained by evaporation of one liter of sea water (as the evaporation proceeds).

Density	Volume	Amounts (in grams) of materials precipitated by evaporation of one liter of sea water. $CaCO_3$	$CaSO_4$*	NaCl
1.0258	1.000	-	-	-
1.0500	0.533	0.064 Range of	-	-
1.0836	0.316	trace precip-	-	-
1.1037	0.245	trace itation	-	-
1.1264	0.190	0.053 of $CaCO_3$	0.443 $CaSO_4$ precip-	-
1.1604	0.145	-	0.445 itated before	-
1.1732	0.131	-	0.144 NaCl begins	-
1.2015	0.112	-	0.127	-
1.2138	0.095	-	0.040	3.261
1.2212	0.064	-	0.117	9.650 Range of
1.2363	0.039	-	0.055	7.896 precip-
1.2570	0.030	-	0.011	2.624 itation
1.2778	0.023	-	-	2.272 of NaCl
1.3069	0.016	-	-	1.404
Total deposit		0.117g.	1.382g.	27.107g.
Amount in last bittern		-	-	2.589g.
Total		0.117g. (i.e. sea water is 0.0117% $CaCO_3$)	1.382g.	29.696g.

*Given by Clarke as $CaSO_4.2H_2O$ but recalculated here as $CaSO_4$.

From:

Petrologic and Geochemical Variations in the Permian Castile Varved Anhydrite Delaware Basin, Texas and New Mexico. By Walter E. Dean, Jr., Ph.D. dissertation, University of New Mexico. 1967, p. 84 a. Data of Usiglio, 1849; modified from Clarke, 1924, p. 220. (Explanatory notes added in the table by D. Wonderly.) Used by permission of the author, W. E. Dean, Jr.